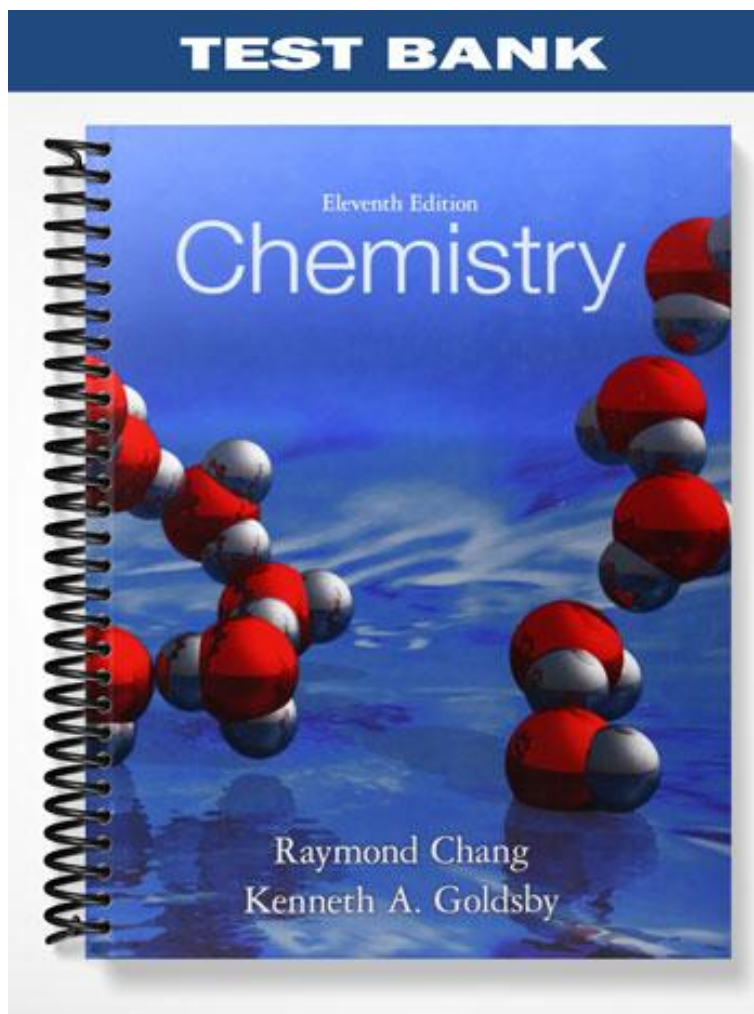


**TEST BANK**



Eleventh Edition

Chemistry

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## Chapter 2: Atoms, Molecules, and Ions

A periodic table is required to work many of the problems in this chapter.

- In a cathode ray tube
  - electrons pass from the anode to the cathode.
  - electrons pass from the cathode to the anode.
  - protons pass from the anode to the cathode.
  - protons pass from the cathode to the anode.Ans: B Category: Medium Section: 2.2
- The scientist who determined the magnitude of the electric charge of the electron was
  - John Dalton.
  - Robert Millikan.
  - J. J. Thomson.
  - Henry Moseley.
  - R. Chang.Ans: B Category: Easy Section: 2.2
- When J. J. Thomson discovered the electron, what physical property of the electron did he measure?
  - its charge,  $e$
  - its charge-to-mass ratio,  $e/m$
  - its temperature,  $T$
  - its mass,  $m$
  - its atomic number,  $Z$Ans: B Category: Easy Section: 2.2
- Which of the following scientists developed the nuclear model of the atom?
  - John Dalton
  - Robert Millikan
  - J. J. Thomson
  - Henry Moseley
  - Ernest RutherfordAns: E Category: Easy Section: 2.2
- Rutherford's experiment with alpha particle scattering by gold foil established that
  - protons are not evenly distributed throughout an atom.
  - electrons have a negative charge.
  - electrons have a positive charge.
  - atoms are made of protons, neutrons, and electrons.
  - protons are 1840 times heavier than electrons.Ans: A Category: Medium Section: 2.2
- Atoms of the same element with different mass numbers are called
  - ions.
  - neutrons.
  - allotropes.
  - chemical families.
  - isotopes.Ans: E Category: Easy Section: 2.3





21. Which one of the following is an ion?  
 A)  $B^{3+}$  B) NaCl C) He D)  $^{14}C$  E) none of the above  
 Ans: A Category: Easy Section: 2.5
22. Which one of the following elements is most likely to form a  $2+$  ion?  
 A) calcium B) carbon C) fluorine D) oxygen E) sodium  
 Ans: A Category: Easy Section: 2.5
23. Which one of the following elements is most likely to form a  $2-$  ion?  
 A) scandium B) selenium C) silicon D) strontium E) iodine  
 Ans: B Category: Easy Section: 2.5
24. A magnesium ion,  $Mg^{2+}$ , has  
 A) 12 protons and 13 electrons. D) 24 protons and 22 electrons.  
 B) 24 protons and 26 electrons. E) 12 protons and 14 electrons.  
 C) 12 protons and 10 electrons.  
 Ans: C Category: Medium Section: 2.5
25. An aluminum ion,  $Al^{3+}$ , has:  
 A) 13 protons and 13 electrons D) 13 protons and 10 electrons  
 B) 27 protons and 24 electrons E) 10 protons and 13 electrons  
 C) 16 protons and 13 electrons  
 Ans: D Category: Medium Section: 2.5
26. An oxide ion,  $O^{2-}$ , has:  
 A) 8 protons and 10 electrons D) 8 protons and 7 electrons  
 B) 10 protons and 8 electrons E) 10 protons and 7 electrons  
 C) 8 protons and 9 electrons  
 Ans: A Category: Medium Section: 2.5
27. A phosphide ion has:  
 A) 10 protons and 13 electrons D) 15 protons and 18 electrons  
 B) 12 protons and 15 electrons E) 18 protons and 21 electrons  
 C) 15 protons and 15 electrons  
 Ans: D Category: Medium Section: 2.5
28. An iron(II) ion has:  
 A) 24 electrons and a charge of  $2+$  D) 28 electrons and a charge of  $2+$   
 B) 24 electrons and a charge of  $2-$  E) 28 electrons and a charge of  $2-$   
 C) 26 electrons and a charge of  $2+$   
 Ans: A Category: Medium Section: 2.5
29. How many protons and electrons are present in one  $Br^-$  ion?  
 A) 35 p, 35 e B) 80 p, 81 e C) 35 p, 34 e D) 35 p, 36 e E) 80 p, 34 e  
 Ans: D Category: Medium Section: 2.5

30. What are the two different ions present in the compound CaS ?  
 A)  $\text{Ca}^+$ ,  $\text{S}^-$  B)  $\text{Ca}^{2-}$ ,  $\text{S}^{2+}$  C)  $\text{Ca}^-$ ,  $\text{S}^+$  D)  $\text{Ca}^{2+}$ ,  $\text{S}^{2-}$  E) Ca, S  
 Ans: D Category: Medium Section: 2.6
31. What are the two different ions present in the compound  $\text{Na}_2\text{S}$ ?  
 A)  $\text{Na}_2^+$ ,  $\text{S}^{2-}$  B)  $\text{Na}^+$ ,  $\text{S}^{2-}$  C)  $\text{Na}^{2+}$ ,  $\text{S}^{2-}$  D)  $\text{Na}^+$ ,  $\text{S}^-$  E)  $\text{Na}^{2+}$ ,  $\text{S}^-$   
 Ans: B Category: Medium Section: 2.6
32. What are the two different ions present in the compound  $\text{Li}_3\text{N}$ ?  
 A)  $\text{Li}^+$ ,  $\text{N}^{3-}$  B)  $\text{Li}_3^+$ ,  $\text{N}^-$  C)  $\text{Li}_3^{3+}$ ,  $\text{N}^{3-}$  D)  $\text{Li}^+$ ,  $\text{N}^-$  E)  $\text{Li}^{3+}$ ,  $\text{N}^{3-}$   
 Ans: A Category: Medium Section: 2.6
33. What are the two different ions present in the compound  $\text{FeCl}_3$ ?  
 A)  $\text{Fe}^{2+}$ ,  $\text{Cl}_3^-$  B)  $\text{Fe}^{3+}$ ,  $\text{Cl}^{3-}$  C)  $\text{Fe}^+$ ,  $\text{Cl}^-$  D)  $\text{Fe}^{3+}$ ,  $\text{Cl}^-$  E)  $\text{Fe}^+$ ,  $\text{Cl}^-$   
 Ans: D Category: Medium Section: 2.6
34. What are the ions present in the compound  $\text{CO}_2$ ?  
 A)  $\text{C}^{4+}$ ,  $2 \text{O}^{2-}$  B)  $\text{C}^{2+}$ ,  $2 \text{O}^-$  C)  $\text{C}^{2+}$ ,  $\text{O}^{2-}$  D)  $\text{C}^{2+}$ ,  $\text{O}_2^{2-}$  E) no ions present  
 Ans: E Category: Medium Section: 2.6
35. What are the ions present in the compound  $\text{CH}_4$ ?  
 A)  $\text{C}^{4+}$ ,  $\text{H}^+$  B)  $\text{C}^{4-}$ ,  $\text{H}^+$  C)  $\text{C}^-$ ,  $\text{H}^+$  D)  $\text{C}^{4-}$ ,  $\text{H}^{4+}$  E) no ions present  
 Ans: E Category: Medium Section: 2.6
36. Which of the following is an example of an empirical formula?  
 A)  $\text{C}_9\text{H}_{12}$  B)  $\text{C}_9\text{H}_{18}\text{Cl}_2$  C)  $\text{C}_6\text{H}_6$  D)  $\text{N}_2\text{O}_4$  E)  $\text{C}_2\text{H}_2\text{O}_2$   
 Ans: B Category: Medium Section: 2.6
37. What is the empirical formula for  $\text{C}_{10}\text{H}_{22}\text{O}_2$ ?  
 A)  $\text{C}_{10}\text{H}_{22}\text{O}_2$  B)  $\text{C}_5\text{H}_{11}\text{O}$  C)  $\text{C}_{20}\text{H}_{44}\text{O}_4$  D)  $\text{C}_2\text{H}_{11}\text{O}$  E)  $\text{C}_5\text{H}_{11}\text{O}_2$   
 Ans: B Category: Medium Section: 2.6
38. What is the empirical formula for  $\text{C}_6\text{H}_{14}\text{O}$ ?  
 A)  $\text{C}_6\text{H}_{14}\text{O}$  B)  $\text{C}_3\text{H}_7\text{O}$  C)  $\text{C}_2\text{H}_7\text{O}$  D)  $\text{C}_{12}\text{H}_{28}\text{O}_2$  E) CHO  
 Ans: A Category: Medium Section: 2.6
39. What is the ion  $\text{ClO}_4^-$  named?  
 A) chloride ion D) perchlorite ion  
 B) chlorite ion E) perchlorate ion  
 C) hypochlorite ion  
 Ans: E Category: Medium Section: 2.7
40. What is the formula for the ionic compound containing calcium ions and nitrate ions?  
 A)  $\text{Ca}_3\text{N}_2$  B)  $\text{Ca}(\text{NO}_3)_2$  C)  $\text{Ca}_2\text{NO}_3$  D)  $\text{Ca}_2\text{NO}_2$  E)  $\text{CaNO}_3$   
 Ans: B Category: Medium Section: 2.7

41. What is the formula for the ionic compound containing calcium ions and oxide ions?  
 A) CaO B) Ca<sub>2</sub>O C) CaO<sub>2</sub> D) Ca<sub>3</sub>O E) CaO<sub>3</sub>  
 Ans: A Category: Medium Section: 2.7
42. What is the formula for the ionic compound containing iron (III) ions and iodide ions?  
 A) FeI B) Fe<sub>2</sub>I C) FeI<sub>2</sub> D) FeI<sub>3</sub> E) Fe<sub>3</sub>I  
 Ans: D Category: Medium Section: 2.7
43. What is the formula for the ionic compound containing sodium ions and nitride ions?  
 A) NaN B) Na<sub>2</sub>N C) NNa<sub>2</sub> D) Na<sub>3</sub>N E) NNa<sub>3</sub>  
 Ans: D Category: Medium Section: 2.7
44. What is the formula for the ionic compound containing barium ions and sulfate ions?  
 A) BaSO<sub>4</sub> B) Ba<sub>2</sub>SO<sub>4</sub> C) BaS D) Ba(SO<sub>4</sub>)<sub>2</sub> E) Ba<sub>3</sub>S<sub>2</sub>  
 Ans: A Category: Medium Section: 2.7
45. What are the two different ions present in the compound Al(NO<sub>3</sub>)<sub>3</sub>?  
 A) Al<sup>3+</sup>, (NO<sub>3</sub>)<sub>3</sub><sup>-</sup> D) Al<sup>3+</sup>, NO<sub>3</sub><sup>3-</sup>  
 B) Al<sup>+</sup>, NO<sub>3</sub><sup>-</sup> E) Al<sup>+</sup>, (NO<sub>3</sub>)<sub>3</sub><sup>-</sup>  
 C) Al<sup>3+</sup>, NO<sub>3</sub><sup>-</sup>  
 Ans: C Category: Medium Section: 2.7
46. What are the two different ions present in the compound NH<sub>4</sub>NO<sub>3</sub>?  
 A) NH<sub>4</sub><sup>-</sup>, NO<sub>3</sub><sup>+</sup> D) NH<sub>4</sub><sup>3+</sup>, NO<sup>4-</sup>  
 B) NH<sub>4</sub><sup>+</sup>, NO<sub>3</sub><sup>-</sup> E) NH<sub>4</sub><sup>+</sup>, NO<sup>3-</sup>  
 C) N<sup>3-</sup>, H<sup>+</sup>, O<sup>2-</sup>  
 Ans: B Category: Medium Section: 2.7
47. Which is the correct formula for iron(II) phosphate?  
 A) Fe<sub>2</sub>PO<sub>4</sub> B) Fe<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> C) Fe<sub>2</sub>PO<sub>3</sub> D) Fe(PO<sub>4</sub>)<sub>2</sub> E) Fe(PO<sub>3</sub>)<sub>2</sub>  
 Ans: B Category: Medium Section: 2.7
48. Which of the following is the formula for hydroiodic acid?  
 A) HIO<sub>4</sub> B) HIO<sub>3</sub> C) HIO<sub>2</sub> D) HIO E) HI  
 Ans: E Category: Medium Section: 2.7
49. The formula for calcium phosphate is  
 A) CaPO<sub>4</sub>. B) Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>. C) Ca<sub>2</sub>(PO<sub>4</sub>)<sub>3</sub>. D) Ca<sub>3</sub>P<sub>2</sub>. E) Ca<sub>3</sub>(PO<sub>3</sub>)<sub>2</sub>.  
 Ans: B Category: Medium Section: 2.7
50. The formula for magnesium sulfate is  
 A) MnS B) MgS C) MnSO<sub>3</sub> D) MgSO<sub>4</sub> E) MgSO<sub>3</sub>  
 Ans: D Category: Medium Section: 2.7

51. The formula for sodium sulfide is  
 A) NaS. B) K<sub>2</sub>S. C) NaS<sub>2</sub>. D) Na<sub>2</sub>S. E) SeS.  
 Ans: D Category: Medium Section: 2.7
52. The name for NH<sub>4</sub>NO<sub>3</sub> is  
 A) ammonium nitrate. D) hydrogen nitrogen oxide.  
 B) ammonium nitrogen trioxide. E) hydrogen nitrate.  
 C) ammonia nitrogen oxide.  
 Ans: A Category: Medium Section: 2.7
53. The name for Ba(OH)<sub>2</sub> is  
 A) barium hydrogen oxide. D) beryllium hydroxide.  
 B) boron hydroxide. E) barium hydroxide.  
 C) barium hydrate.  
 Ans: E Category: Medium Section: 2.7
54. The name for KHCO<sub>3</sub> is  
 A) calcium bicarbonate. D) calcium hydrogen carbon trioxide.  
 B) calcium carbonate. E) potassium hydrogen carbonate.  
 C) potassium carbonate.  
 Ans: E Category: Medium Section: 2.7
55. The name for CuSO<sub>4</sub>·5H<sub>2</sub>O is  
 A) copper sulfate acid. D) copper(II) sulfate pentahydrate.  
 B) copper sulfate pentahydrate. E) copper(V) sulfate hydrate.  
 C) copper(II) sulfate acid.  
 Ans: D Category: Medium Section: 2.7
56. Give the formula for cobalt(II) chlorate dihydrate  
 A) CoCl<sub>2</sub>·2H<sub>2</sub>O D) Co(ClO<sub>3</sub>)<sub>2</sub>·2H<sub>2</sub>O  
 B) CoClO<sub>3</sub>(H<sub>2</sub>O)<sub>2</sub> E) Co<sub>2</sub>(ClO<sub>3</sub>)<sub>3</sub>·2H<sub>2</sub>O  
 C) Co(ClO<sub>3</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>2</sub>  
 Ans: D Category: Medium Section: 2.7
57. Name the compound Co(NO<sub>3</sub>)<sub>2</sub>  
 A) Cobalt (I) nitrate D) Cobalt nitrite  
 B) Cobalt (II) nitrate E) Cobalt (II) nitride  
 C) Cobalt (I) nitride  
 Ans: B Category: Medium Section: 2.7
58. Name the compound CuSO<sub>4</sub>  
 A) Copper (I) sulfate D) Copper (II) sulfate  
 B) Copper (I) sulfite E) Copper (IV) sulfate  
 C) Copper (II) sulfite  
 Ans: D Category: Medium Section: 2.7







77. The straight chain hydrocarbon that contains six carbon atoms is  
A) propane B) butane C) pentane D) hexane E) heptane  
Ans: D Category: Medium Section: 2.8
78. What is the law of conservation of mass?  
Ans: Matter can be neither created nor destroyed.  
Category: Easy Section: 2.1
79. Describe the contributions of Marie Curie.  
Ans: (note that answers will vary) Marie Curie discovered two new elements, and is one of three people to win two Nobel Prizes. She also suggested the term “radioactivity” to describe the spontaneous emission of particles and/or radiation.  
Category: Easy Section: 2.2
80. What are the three types of radiation that can be produced by the decay of radioactive substances like uranium?  
Ans: Alpha, beta, and gamma radiation  
Category: Easy Section: 2.2
81. Marie Curie suggested the name “radioactivity” to describe the spontaneous emission of particles and/or radiation.  
Ans: True Category: Easy Section: 2.2
82. Using a cathode ray tube, J. J. Thomson determined the magnitude of the electric charge on the electron.  
Ans: False Category: Easy Section: 2.2
83. When a beam of alpha particles passes between two electrically charged plates, the beam is deflected toward the positive plate.  
Ans: False Category: Medium Section: 2.2
84. The proton is about 1840 times heavier than the electron.  
Ans: True Category: Easy Section: 2.2
85. How many electrons, protons, and neutrons does an iron-55 atom have?  
Ans: 26 electrons, 26 protons, and 29 neutrons  
Category: Medium Section: 2.3
86. How many protons are there in one atom of nickel?  
Ans: 28  
Category: Easy Section: 2.3
87. How many protons are there in one atom of magnesium?  
Ans: 12  
Category: Easy Section: 2.3

88. How many protons are there in one atom of xenon?

Ans: 54

Category: Easy Section: 2.3

89. Almost all the mass of an atom is concentrated in the nucleus.

Ans: True Category: Easy Section: 2.3

90. The atomic number is equal to the number of protons in the nucleus of each atom of an element.

Ans: True Category: Easy Section: 2.3

91. The number of neutrons in all atoms of an element is the same.

Ans: False Category: Medium Section: 2.3

92. How many protons are there in one atom of uranium?

Ans: 92

Category: Easy Section: 2.3

93. What are *isotopes*?

Ans: Atoms of the same element that have the same atomic number but different mass numbers.

Category: Easy Section: 2.3

94. The table below describes four atoms.

	Atom A	Atom B	Atom C	Atom D
Number of protons	79	80	80	79
Number of neutrons	118	120	118	120
Number of electrons	79	80	80	79

Which atoms represent the same element?

Ans: Atoms A and D represent the same element, and atoms B and C represent the same element.

Category: Medium Section: 2.3

95. Consider a neutral atom of the following isotope of sulfur:



How many electrons, protons, and neutrons does the atom contain?

Ans: 16 electrons, 16 protons, and 18 neutrons

Category: Medium Section: 2.3



101. Use the periodic table above to show where the halogen elements are located.  
Ans: Group 7A or Group 17  
Category: Easy Section: 2.4
102. Use the periodic table above to show where the noble gases are located.  
Ans: Group 8A or Group 18  
Category: Easy Section: 2.4
103. The elements known as the halogens are useful as disinfectants. Name two halogens.  
Ans: (two of these) fluorine, chlorine, bromine, iodine  
Category: Easy Section: 2.4
104. Define the term *molecule*.  
Ans: A molecule is an aggregate of at least two atoms in a definite arrangement held together by chemical forces.  
Category: Easy Section: 2.5
105. What are the seven elements that naturally occur as diatomic molecules?  
Ans: Hydrogen, nitrogen, oxygen, fluorine, chlorine, bromine, iodine  
Category: Medium Section: 2.5
106. Define *ion*.  
Ans: An ion is an atom or group of atoms that has a net positive or negative charge.  
Category: Easy Section: 2.5
107. A molecule of antifreeze, ethylene glycol, has the formula  $C_2H_4(OH)_2$ . How many atoms are there in one molecule of antifreeze?  
Ans: 10  
Category: Easy Section: 2.5
108. How many carbon atoms are in one molecule of  $CH_3(CH_2)_3CH_3$ ?  
Ans: 5  
Category: Easy Section: 2.5
109. How many hydrogen atoms are in one molecule of  $CH_3(CH_2)_3CH_3$ ?  
Ans: 12  
Category: Easy Section: 2.5
110. The formula for isopropyl alcohol is sometimes written as  $(CH_3)_2CHOH$  to better indicate how the atoms are connected. How many hydrogen atoms would be contained in 3 dozen isopropyl alcohol molecules?  
Ans: 288  
Category: Medium Section: 2.5

111. Define *allotrope*.

Ans: An allotrope is one of the two or more distinct forms of an element.

Category: Easy Section: 2.6

112. An empirical formula tell us which elements are present in a compound and gives us the simplest, whole-number ratio of the atoms of these elements in the compound.

Ans: True Category: Easy Section: 2.6

113. Give the formula for potassium oxide.

Ans:  $K_2O$

Category: Medium Section: 2.7

114. Give the formula for magnesium chloride.

Ans:  $MgCl_2$

Category: Medium Section: 2.7

115. Give the formula for carbon disulfide.

Ans:  $CS_2$

Category: Medium Section: 2.7

116. Give the formula for potassium hydroxide.

Ans:  $KOH$

Category: Medium Section: 2.7

117. Give the formula for nickel(II) sulfite.

Ans:  $NiSO_3$

Category: Medium Section: 2.7

118. Name the following binary compound:  $FeS$ .

Ans: iron(II) sulfide or ferrous sulfide

Category: Medium Section: 2.7

119. Name the following binary compound:  $NaH$ .

Ans: sodium hydride

Category: Medium Section: 2.7

120. Name the following binary compound:  $MnCl_2$ .

Ans: manganese(II) chloride or manganous chloride

Category: Medium Section: 2.7

121. Name the following binary compound:  $Fe_2O_3$ .

Ans: iron(III) oxide (or ferric oxide)

Category: Medium Section: 2.7

122. Name the following compound:  $\text{CuCO}_3$ .  
Ans: copper(II) carbonate or cupric carbonate  
Category: Medium Section: 2.7
123. Name the following compound:  $\text{K}_3\text{PO}_4$ .  
Ans: potassium phosphate  
Category: Medium Section: 2.7
124. Name the following compound:  $\text{Al}(\text{NO}_2)_2$ .  
Ans: aluminum nitrite  
Category: Medium Section: 2.7
125. Name the following compound:  $\text{Cl}_2\text{O}_7$ .  
Ans: dichlorine heptoxide  
Category: Medium Section: 2.7
126. Give the formula of magnesium nitrate.  
Ans:  $\text{Mg}(\text{NO}_3)_2$   
Category: Medium Section: 2.7
127. Give the formula of calcium phosphate.  
Ans:  $\text{Ca}_3(\text{PO}_4)_2$   
Category: Medium Section: 2.7
128. Give the formula of iron(II) phosphate.  
Ans:  $\text{Fe}_3(\text{PO}_4)_2$   
Category: Medium Section: 2.7
129. Give the formula of copper(II) bromide.  
Ans:  $\text{CuBr}_2$   
Category: Medium Section: 2.7
130. Give the formula of ammonium sulfate.  
Ans:  $(\text{NH}_4)_2\text{SO}_4$   
Category: Medium Section: 2.7
131. Give the formula of hydrochloric acid.  
Ans:  $\text{HCl}$   
Category: Medium Section: 2.7
132. Give the formula of carbonic acid.  
Ans:  $\text{H}_2\text{CO}_3$   
Category: Medium Section: 2.7



133. Give the formula of nitrous acid.  
Ans:  $\text{HNO}_2$   
Category: Medium Section: 2.7
134. Give the formula of sulfuric acid.  
Ans:  $\text{H}_2\text{SO}_4$   
Category: Medium Section: 2.7
135. Name the following: HF.  
Ans: hydrofluoric acid  
Category: Medium Section: 2.7
136. Name the following  $\text{H}_3\text{PO}_3$   
Ans: phosphorous acid  
Category: Medium Section: 2.7
137. Write the formula of ammonia.  
Ans:  $\text{NH}_3$   
Category: Medium Section: 2.7
138. Write the formula of lead(II) chloride.  
Ans:  $\text{PbCl}_2$   
Category: Medium Section: 2.7
139. Write the formula of calcium carbonate.  
Ans:  $\text{CaCO}_3$   
Category: Medium Section: 2.7
140. Write the formula of an anion that contains a metal.  
Ans:  $\text{CrO}_4^{2-}$  or  $\text{Cr}_2\text{O}_7^{2-}$  or  $\text{MnO}_4^-$   
Category: Medium Section: 2.7
141. Write the formula of a cation that contains a nonmetal.  
Ans:  $\text{NH}_4^+$   
Category: Medium Section: 2.7
142. Give an example of an anion that contains a metal and write the name.  
Ans: chromate or dichromate or permanganate  
Category: Medium Section: 2.7
143. What is the nitride ion, nitrate ion, and nitrite ion, in that order?  
Ans:  $\text{N}^{3-}$ ,  $\text{NO}_3^-$ , and  $\text{NO}_2^-$   
Category: Medium Section: 2.7

144. What is the sulfide ion, sulfate ion, and sulfite ion, in that order?  
Ans:  $S^{2-}$ ,  $SO_4^{2-}$ ,  $SO_3^{2-}$   
Category: Medium Section: 2.7
145. What is the chloride ion, chlorate ion, and perchlorate ion, in that order?  
Ans:  $Cl^-$ ,  $ClO_3^-$ , and  $ClO_4^-$   
Category: Medium Section: 2.7
146. What is chloric acid, chlorous acid, and hypochlorous acid, in that order?  
Ans:  $HClO_3$ ,  $HClO_2$ ,  $HClO$   
Category: Medium Section: 2.7
147. What is ammonia and ammonium ion, in that order?  
Ans:  $NH_3$ ,  $NH_4^+$   
Category: Medium Section: 2.7
148. What is the formula for dinitrogen monoxide?  
Ans:  $N_2O$   
Category: Medium Section: 2.7
149. What is the formula for dibromine heptoxide?  
Ans:  $Br_2O_7$   
Category: Medium Section: 2.7
150. What is the formula for xenon difluoride?  
Ans:  $XeF_2$   
Category: Medium Section: 2.7
151. What is the formula for xenon hexafluoride?  
Ans:  $XeF_6$   
Category: Medium Section: 2.7
152. What is the formula for the compound hydrogen peroxide  
Ans:  $H_2O_2$   
Category: Medium Section: 2.7
153. Name the compound  $CH_3CH_2OH$   
Ans: Ethanol  
Category: Medium Section: 2.8
154. Name the compound  $CH_3CH_2NH_2$   
Ans: Ethylamine  
Category: Medium Section: 2.8

155. What is the formula for octane?

Ans:  $C_8H_{18}$

Category: Medium Section: 2.8

156. What is the formula for nonane?

Ans:  $C_9H_{20}$

Category: Medium Section: 2.8