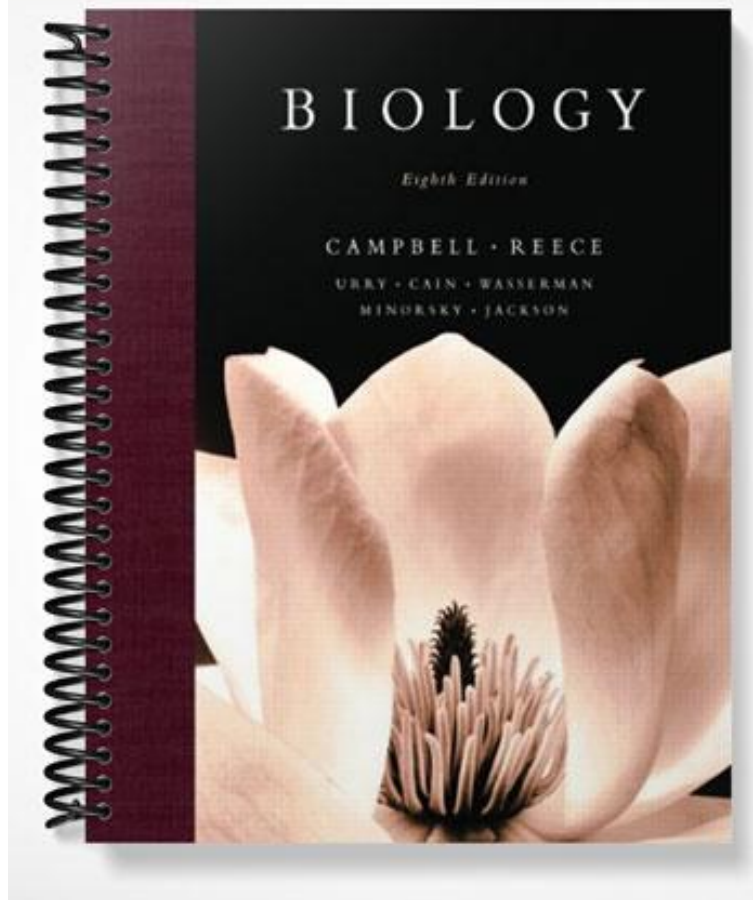


TEST BANK



BIOLOGY

Eighth Edition

CAMPBELL · REECE

URRY · CAIN · WASSERMAN
MINORSKY · JACKSON

The Chemical Context of Life

Multiple-Choice Questions

1) About 25 of the 92 natural elements are known to be essential to life. Which four of these 25 elements make up approximately 96% of living matter?

- A) carbon, sodium, chlorine, nitrogen
- B) carbon, sulfur, phosphorus, hydrogen
- C) oxygen, hydrogen, calcium, sodium
- D) carbon, hydrogen, nitrogen, oxygen
- E) carbon, oxygen, sulfur, calcium

Answer: D

Topic: Concept 2.1

Skill: Knowledge/Comprehension

2) Trace elements are those required by an organism in only minute quantities. Which of the following is a trace element that is required by humans and other vertebrates?

- A) nitrogen
- B) calcium
- C) iodine
- D) sodium
- E) phosphorus

Answer: C

Topic: Concept 2.1

Skill: Knowledge/Comprehension

3) Three or four of the following statements are true and correct. Which one, if any, is false? If all the statements are true, choose answer E.

- A) Carbon, hydrogen, oxygen, and nitrogen make up approximately 96% of living matter.
- B) The trace element iodine is required only in very small quantities by vertebrates.
- C) Virtually all organisms require the same elements in the same quantities.
- D) Iron is an example of an element needed by all organisms.
- E) All of the other statements are true and correct.

Answer: C

Topic: Concept 2.1

Skill: Knowledge/Comprehension

4) Which of the following statements is false?

- A) Atoms of the various elements differ in their number of subatomic particles.
- B) All atoms of a particular element have the same number of protons in their nuclei.
- C) The neutrons and protons present in the nucleus of an atom are almost identical in mass; each has a mass of about 1 dalton.

- D) An atom is the smallest unit of an element that still retains the properties of the element.
E) Protons and electrons are electrically charged particles. Protons have one unit of positive charge, and electrons have one unit of negative charge.

Answer: E

Topic: Concept 2.2

Skill: Knowledge/Comprehension

5) Each element is unique and different from other elements because of the number of protons in the nuclei of its atoms. Which of the following indicates the number of protons in an atom's nucleus?

- A) atomic mass
- B) atomic weight
- C) atomic number
- D) mass weight
- E) mass number

Answer: C

Topic: Concept 2.2

Skill: Knowledge/Comprehension

6) The mass number of an element can be easily approximated by adding together the number of _____ in an atom of that element.

- A) protons and neutrons
- B) energy levels
- C) protons and electrons
- D) neutrons and electrons
- E) isotopes

Answer: A

Topic: Concept 2.2

Skill: Knowledge/Comprehension

7) What is the approximate atomic mass of an atom with 16 neutrons, 15 protons, and 15 electrons?

- A) 15 daltons
- B) 16 daltons
- C) 30 daltons
- D) 31 daltons
- E) 46 daltons

Answer: D

Topic: Concept 2.2

Skill: Knowledge/Comprehension

8) Oxygen has an atomic number of 8 and a mass number of 16. Thus, the atomic mass of an oxygen atom is

- A) exactly 8 grams.

- B) exactly 8 daltons.
- C) approximately 16 grams.
- D) approximately 16 daltons.
- E) 24 amu (atomic mass units).

Answer: D

Topic: Concept 2.2

Skill: Knowledge/Comprehension

9) The nucleus of a nitrogen atom contains 7 neutrons and 7 protons. Which of the following is a correct statement concerning nitrogen?

- A) The nitrogen atom has a mass number of approximately 7 daltons and an atomic mass of 14.
- B) The nitrogen atom has a mass number of approximately 14 daltons and an atomic mass of 7.
- C) The nitrogen atom has a mass number of 14 and an atomic mass of 7 grams.
- D) The nitrogen atom has a mass number of 7 grams and an atomic number of 14.
- E) The nitrogen atom has a mass number of 14 and an atomic mass of approximately 14 daltons.

Answer: E

Topic: Concept 2.2

Skill: Knowledge/Comprehension

10) Calcium has an atomic number of 20 and an atomic mass of 40. Therefore, a calcium atom must have

- A) 20 protons.
- B) 40 electrons.
- C) 40 neutrons.
- D) A and B only
- E) A, B, and C

Answer: A

Topic: Concept 2.2

Skill: Knowledge/Comprehension

11) An atom with an atomic number of 9 and a mass number of 19 would have an atomic mass of approximately

- A) 9 daltons.
- B) 9 grams.
- C) 10 daltons.
- D) 20 grams.
- E) 19 daltons.

Answer: E

Topic: Concept 2.2

Skill: Knowledge/Comprehension

12) Different atomic forms of an element contain the same number of protons but a different number of neutrons. What are these different atomic forms called?

- A) ions

- B) isotopes
- C) neutronic atoms
- D) isomers
- E) radioactive atoms

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

13) How do isotopes of the same element differ from each other?

- A) number of protons
- B) number of electrons
- C) number of neutrons
- D) valence electron distribution
- E) amount of radioactivity

Answer: C

Topic: Concept 2.2

Skill: Knowledge/Comprehension

14) Which of the following best describes the relationship between the atoms described below?



- A) They are isomers.
- B) They are polymers.
- C) They are isotopes.
- D) They contain 1 and 3 protons, respectively.
- E) They each contain 1 neutron.

Answer: C

Topic: Concept 2.2

Skill: Knowledge/Comprehension

15) Which of the following best describes the relationship between the atoms described below?



- A) They contain 31 and 32 electrons, respectively.
- B) They are both phosphorus cations.
- C) They are both phosphorus anions.
- D) They are both isotopes of phosphorus.
- E) They contain 31 and 32 protons, respectively.

Answer: D

Topic: Concept 2.2

Skill: Knowledge/Comprehension

16) One difference between carbon-12 ($^{12}_6\text{C}$) and carbon-14 ($^{14}_6\text{C}$) is that carbon-14 has

- A) two more protons than carbon-12.
- B) two more electrons than carbon-12.
- C) two more neutrons than carbon-12.
- D) A and C only
- E) B and C only

Answer: C

Topic: Concept 2.2

Skill: Knowledge/Comprehension

17) ^3_1H is a radioactive isotope of hydrogen. One difference between hydrogen-1 (^1_1H) and hydrogen-3 (^3_1H) is that hydrogen-3 has

- A) one more neutron and one more proton than hydrogen-1.
- B) one more proton and one more electron than hydrogen-1.
- C) one more electron and one more neutron than hydrogen-1.
- D) two more neutrons than hydrogen-1.
- E) two more protons than hydrogen-1.

Answer: D

Topic: Concept 2.2

Skill: Knowledge/Comprehension

18) The atomic number of carbon is 6. Carbon-14 is heavier than carbon-12 because the atomic nucleus of carbon-14 contains _____ neutrons.

- A) 6
- B) 7
- C) 8
- D) 12
- E) 14

Answer: C

Topic: Concept 2.2

Skill: Knowledge/Comprehension

19) Electrons exist only at fixed levels of potential energy. However, if an atom absorbs sufficient energy, a possible result is that

- A) an electron may move to an electron shell farther out from the nucleus.
- B) an electron may move to an electron shell closer to the nucleus.
- C) the atom may become a radioactive isotope.
- D) the atom would become a positively charged ion, or cation.
- E) the atom would become a negatively charged ion, or anion.

Answer: A

Topic: Concept 2.2
Skill: Knowledge/Comprehension

20) The atomic number of neon is 10. Therefore, which of the following is correct about an atom of neon?

- A) It has 8 electrons in its outer electron shell.
- B) It is inert.
- C) It has an atomic mass of 10 daltons.
- D) A and B only
- E) A, B, and C are correct.

Answer: D

Topic: Concept 2.2
Skill: Knowledge/Comprehension

21) From its atomic number of 15, it is possible to predict that the phosphorus atom has

- A) 15 neutrons.
- B) 15 protons.
- C) 15 electrons.
- D) 8 electrons in its outermost electron shell.
- E) B and C only

Answer: E

Topic: Concept 2.2
Skill: Knowledge/Comprehension

Please refer to Figure 2.1 to answer the following questions.

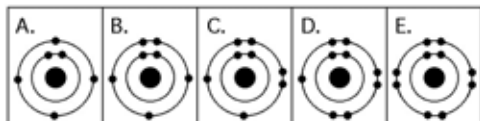


Figure 2.1

22) Which drawing depicts the electron configuration of neon (${}^{20}_{10}\text{Ne}$)?

Answer: E

Topic: Concept 2.2
Skill: Knowledge/Comprehension

23) Which drawing depicts the electron configuration of oxygen (${}^{16}_8\text{O}$)?

Answer: C

Topic: Concept 2.2
Skill: Knowledge/Comprehension

24) Which drawing depicts the electron configuration of nitrogen (${}^{14}_7\text{N}$)?

Answer: B
Topic: Concept 2.2
Skill: Knowledge/Comprehension

25) Which drawing is of an atom with the atomic number of 6?

Answer: A
Topic: Concept 2.2
Skill: Knowledge/Comprehension

26) Which drawing depicts an atom that is inert or chemically unreactive?

Answer: E
Topic: Concept 2.2
Skill: Knowledge/Comprehension

27) Which drawing depicts an atom with a valence of 3?

Answer: B
Topic: Concept 2.2
Skill: Knowledge/Comprehension

28) Which drawing depicts an atom with a valence of 2?

Answer: C
Topic: Concept 2.2
Skill: Knowledge/Comprehension

29) Atoms whose outer electron shells contain eight electrons tend to

- A) form ionic bonds in aqueous solutions.
- B) form covalent bonds in aqueous solutions.
- C) be stable and chemically nonreactive, or inert.
- D) be unstable and chemically very reactive.
- E) be isotopes and very radioactive.

Answer: C
Topic: Concept 2.2
Skill: Knowledge/Comprehension

Use the information extracted from the periodic table in Figure 2.2 to answer the following questions.

Atomic mass →	12	16	1	14	32	31
	C	O	H	N	S	P
Atomic number →	6	8	1	7	16	15

Figure 2.2

30) How many electrons does nitrogen have in its valence shell?

- A) 2
- B) 5
- C) 7
- D) 8
- E) 14

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

31) How many electrons does phosphorus have in its valence shell?

- A) 1
- B) 2
- C) 3
- D) 4
- E) 5

Answer: E

Topic: Concept 2.2

Skill: Knowledge/Comprehension

32) How many neutrons are present in the nucleus of a phosphorus atom?

- A) 8
- B) 15
- C) 16
- D) 31
- E) 46

Answer: C

Topic: Concept 2.2

Skill: Knowledge/Comprehension

33) How many electrons does an atom of sulfur have in its valence shell?

- A) 4
- B) 6
- C) 8
- D) 16
- E) 32

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

34) Based on electron configuration, which of these elements would exhibit chemical behavior most like that of oxygen?

- A) carbon

- B) hydrogen
- C) nitrogen
- D) sulfur
- E) phosphorus

Answer: D

Topic: Concept 2.2

Skill: Application/Analysis

35) How many electrons would be expected in the outermost electron shell of an atom with atomic number 12?

- A) 1
- B) 2
- C) 4
- D) 6
- E) 8

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

36) The atomic number of each atom is given to the left of each of the elements below. Which of the atoms has the same valence as carbon (${}^{12}_6\text{C}$)?

- A) 7 nitrogen
- B) 9 fluorine
- C) 10 neon
- D) 12 magnesium
- E) 14 silicon

Answer: E

Topic: Concept 2.2

Skill: Application/Analysis

37) What is the valence of an atom with six electrons in its outer electron shell?

- A) 1
- B) 2
- C) 3
- D) 4
- E) 5

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

38) Fluorine has an atomic number of 9 and a mass number of 19. How many electrons are needed to complete the valence shell of a fluorine atom?

- A) 1

- B) 3
- C) 5
- D) 7
- E) 9

Answer: A

Topic: Concept 2.2

Skill: Knowledge/Comprehension

39) What is the maximum number of electrons in the $1s$ orbital of an atom?

- A) 1
- B) 2
- C) 3
- D) 4
- E) 5

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

40)

What is the maximum number of electrons in a $2p$ orbital of an atom?

- A) 1
- B) 2
- C) 3
- D) 4
- E) 5

Answer: B

Topic: Concept 2.2

Skill: Knowledge/Comprehension

41) A covalent chemical bond is one in which

- A) electrons are removed from one atom and transferred to another atom so that the two atoms become oppositely charged.
- B) protons and neutrons are shared by two atoms so as to satisfy the requirements of both atoms.
- C) outer-shell electrons of two atoms are shared so as to satisfactorily fill the outer electron shells of both atoms.
- D) outer-shell electrons of one atom are transferred to the inner electron shells of another atom.
- E) the inner-shell electrons of one atom are transferred to the outer shell of another atom.

Answer: C

Topic: Concept 2.3

Skill: Knowledge/Comprehension

42) If an atom of sulfur (atomic number 16) were allowed to react with atoms of hydrogen (atomic number 1), which of the molecules below would be formed?

- A) S — H
 B) H — S — H
 C) H — S — H
 |
 H
 D) H
 H
 |
 H — S — H
 |
 H
 E) H=S=H

Answer: B

Topic: Concept 2.3

Skill: Application/Analysis

43) What is the maximum number of covalent bonds an element with atomic number 8 can make with hydrogen?

- A) 1
 B) 2
 C) 3
 D) 4
 E) 6

Answer: B

Topic: Concept 2.3

Skill: Knowledge/Comprehension

44) A molecule of carbon dioxide (CO₂) is formed when one atom of carbon (atomic number 6) is covalently bonded with two atoms of oxygen (atomic number 8). What is the total number of electrons that must be shared between the carbon atom and the oxygen atoms in order to complete the outer electron shell of all three atoms?

- A) 1
 B) 2
 C) 3
 D) 4
 E) 5

Answer: D

Topic: Concept 2.3

Skill: Application/Analysis

45) Nitrogen (N) is much more electronegative than hydrogen (H). Which of the following statements is correct about the atoms in ammonia (NH₃)?

- A) Each hydrogen atom has a partial positive charge.
 B) The nitrogen atom has a strong positive charge.
 C) Each hydrogen atom has a slight negative charge.
 D) The nitrogen atom has a partial positive charge.

E) There are covalent bonds between the hydrogen atoms.

Answer: A

Topic: Concept 2.3

Skill: Knowledge/Comprehension

46) When two atoms are equally electronegative, they will interact to form

- A) equal numbers of isotopes.
- B) ions.
- C) polar covalent bonds.
- D) nonpolar covalent bonds.
- E) ionic bonds.

Answer: D

Topic: Concept 2.3

Skill: Knowledge/Comprehension

47) What results from an unequal sharing of electrons between atoms?

- A) a nonpolar covalent bond
- B) a polar covalent bond
- C) an ionic bond
- D) a hydrogen bond
- E) a hydrophobic interaction

Answer: B

Topic: Concept 2.3

Skill: Knowledge/Comprehension

48) A covalent bond is likely to be polar when

- A) one of the atoms sharing electrons is much more electronegative than the other atom.
- B) the two atoms sharing electrons are equally electronegative.
- C) the two atoms sharing electrons are of the same element.
- D) it is between two atoms that are both very strong electron acceptors.
- E) the two atoms sharing electrons are different elements.

Answer: A

Topic: Concept 2.3

Skill: Knowledge/Comprehension

49) Which of the following molecules contains the strongest polar covalent bond?

- A) H₂
- B) O₂
- C) CO₂
- D) H₂O
- E) CH₄

Answer: D

Topic: Concept 2.3
Skill: Knowledge/Comprehension

The following questions refer to Figure 2.3.

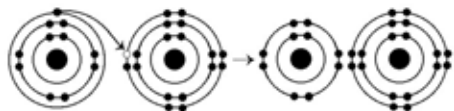


Figure 2.3

50) What results from the chemical reaction illustrated in Figure 2.3?

- A) a cation with a net charge of +1
- B) a cation with a net charge of -1
- C) an anion with a net charge of +1
- D) an anion with a net charge of -1
- E) A and D

Answer: E

Topic: Concept 2.3
Skill: Knowledge/Comprehension

51) What is the atomic number of the cation formed in the reaction illustrated in Figure 2.3?

- A) 1
- B) 8
- C) 10
- D) 11
- E) 16

Answer: D

Topic: Concept 2.3
Skill: Application/Analysis

52) The ionic bond of sodium chloride is formed when

- A) chlorine gains an electron from sodium.
- B) sodium and chlorine share an electron pair.
- C) sodium and chlorine both lose electrons from their outer valence shells.
- D) sodium gains an electron from chlorine.
- E) chlorine gains a proton from sodium.

Answer: A

Topic: Concept 2.3
Skill: Knowledge/Comprehension

53) What is the difference between covalent bonds and ionic bonds?

- A) Covalent bonds involve the sharing of protons between atoms, and ionic bonds involve the sharing of electrons between atoms.

- B) Covalent bonds involve the sharing of neutrons between atoms, and ionic bonds involve the sharing of electrons between atoms.
- C) Covalent bonds involve the sharing of electrons between atoms, and ionic bonds involve the electrical attraction between atoms.
- D) Covalent bonds involve the sharing of protons between atoms, and ionic bonds involve the sharing of neutrons between atoms.
- E) Covalent bonds involve the transfer of electrons between atoms, and ionic bonds involve the sharing of neutrons between atoms.

Answer: C

Topic: Concept 2.3

Skill: Knowledge/Comprehension

54)

In ammonium chloride salt (NH_4Cl) the anion is a single chloride ion, Cl^- . What is the cation of NH_4Cl ?

- A) N, with a charge of +3
- B) H, with a charge of +1
- C) H_2 with a charge of +4
- D) NH_4 with a charge of +1
- E) NH_4 with a charge of +4

Answer: D

Topic: Concept 2.3

Skill: Knowledge/Comprehension

55) The atomic number of chlorine is 17. The atomic number of magnesium is 12. What is the formula for magnesium chloride?

- A) MgCl
- B) MgCl_2
- C) Mg_2Cl
- D) Mg_2Cl_2
- E) MgCl_3

Answer: B

Topic: Concept 2.3

Skill: Application/Analysis

56) Which of the following results from a transfer of electron(s) between atoms?

- A) nonpolar covalent bond
- B) polar covalent bond
- C) ionic bond
- D) hydrogen bond
- E) hydrophobic interaction

Answer: C

Topic: Concept 2.3

Skill: Knowledge/Comprehension

57)

Which of the following explains most specifically the attraction of water molecules to one another?

- A) nonpolar covalent bond
- B) polar covalent bond
- C) ionic bond
- D) hydrogen bond
- E) hydrophobic interaction

Answer: D

Topic: Concept 2.3

Skill: Knowledge/Comprehension

58) Van der Waals interactions result when

- A) hybrid orbitals overlap.
- B) electrons are not symmetrically distributed in a molecule.
- C) molecules held by ionic bonds react with water.
- D) two polar covalent bonds react.
- E) a hydrogen atom loses an electron.

Answer: B

Topic: Concept 2.3

Skill: Knowledge/Comprehension

59) A van der Waals interaction is the weak attraction between

- A) the electrons of one molecule and the electrons of a nearby molecule.
- B) the nucleus of one molecule and the electrons of a nearby molecule.
- C) a polar molecule and a nearby nonpolar molecule.
- D) a polar molecule and a nearby molecule that is also polar.
- E) a nonpolar molecule and a nearby molecule that is also nonpolar.

Answer: B

Topic: Concept 2.3

Skill: Knowledge/Comprehension

60) Which of the following is not considered to be a weak molecular interaction?

- A) a covalent bond
- B) a van der Waals interaction
- C) an ionic bond in the presence of water
- D) a hydrogen bond
- E) A and B only

Answer: A

Topic: Concept 2.3

Skill: Knowledge/Comprehension

61) Which of the following would be regarded as compounds?

A) H_2

B) H_2O

C) O_2

D) CH_4

E) B and D, but not A and C

Answer: E

Topic: Concept 2.3

Skill: Application/Analysis

62) Sometimes atoms form molecules by sharing two pairs of valence electrons. When this occurs, the atoms are said to be joined by

A) a double covalent bond.

B) an electronegative bond.

C) a hydrogen bond.

D) a protonic bond.

E) a complex bond.

Answer: A

Topic: Concept 2.3

Skill: Knowledge/Comprehension

Refer to the following figure to answer the following questions.



63) The molecule shown here could be described in chemical symbols as

A) CH_4 .

B) H_2O .

C) C_2H_3 .

D) C_4H_4 .

E) CH_2O .

Answer: A

Topic: Concept 2.3

Skill: Application/Analysis

64) The molecule shown here is the simplest of organic compounds. It is called

A) a carbohydrate.

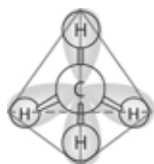
- B) carbon dioxide.
- C) methane.
- D) carbonic hydrate.
- E) methyl carbonate.

Answer: C

Topic: Concept 2.3

Skill: Application/Analysis

Refer to the following figure to answer the following questions.



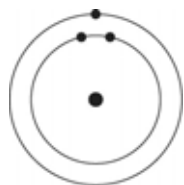
- 65) In the methane molecule shown here, bonds have formed that include both the s orbital valence electrons of the hydrogen atoms and the p orbital valence electrons of the carbon. The electrons in these bonds are said to have
- A) double orbitals.
 - B) tetrahedral orbitals.
 - C) complex orbitals.
 - D) hybrid orbitals.
 - E) reduced orbitals.

Answer: D

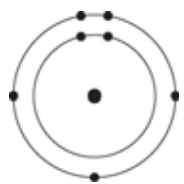
Topic: Concept 2.3

Skill: Application/Analysis

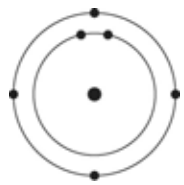
66) Which one of the atoms shown would be most likely to form a cation with a charge of +1?
A)



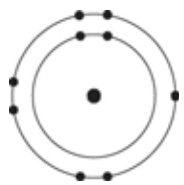
B)



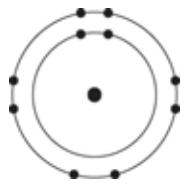
C)



D)



E)

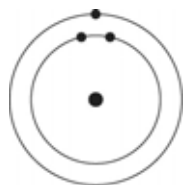


Answer: A

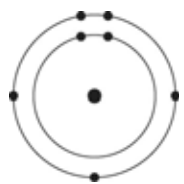
Topic: Concept 2.3

Skill: Application/Analysis

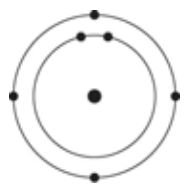
67) Which one of the atoms shown would be most likely to form an anion with a charge of -1?
A)



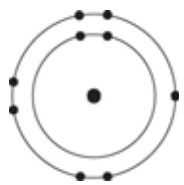
B)



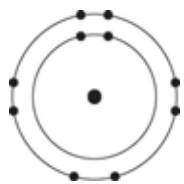
C)



D)



E)



Answer: D

Topic: Concept 2.3

Skill: Application/Analysis

68) Which of the following pairs of atoms would be most likely to form a covalent bond?

A)



B)



C)



D)



E)



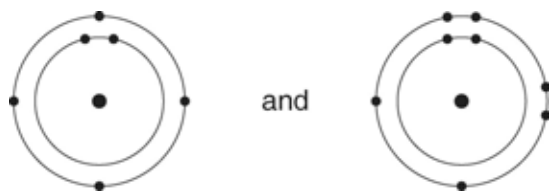
Answer: A

Topic: Concept 2.3

Skill: Application/Analysis

69) Which of the following pairs of atoms would be most likely to form an ionic bond?

A)



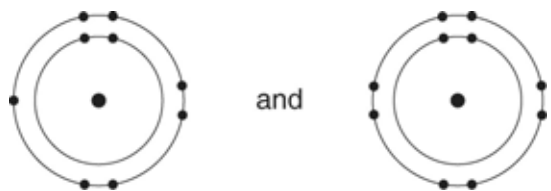
B)



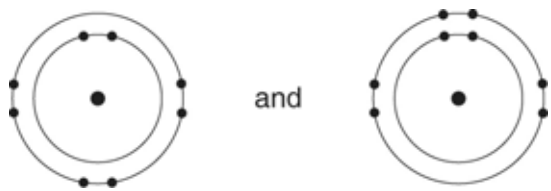
C)



D)



E)



Answer: B

Topic: Concept 2.3

Skill: Application/Analysis

70) The hybrid orbitals in a molecule of methane are oriented

A) toward the corners of a tetrahedron centered on the carbon atom.

B) toward the corners of a cube centered on the carbon atom.

- C) toward the corners of a triangle centered on the carbon atom.
- D) toward the corners of a rectangle centered on the carbon atom.
- E) toward the edges of an oval centered on the carbon atom.

Answer: A

Topic: Concept 2.3

Skill: Application/Analysis

71) Which of the following is true for this reaction? $3 \text{H}_2 + \text{N}_2 \rightleftharpoons 2 \text{NH}_3$

- A) The reaction is nonreversible.
- B) Hydrogen and nitrogen are the reactants of the reverse reaction.
- C) Hydrogen and nitrogen are the products of the forward reaction.
- D) Ammonia is being formed and decomposed.
- E) Hydrogen and nitrogen are being decomposed.

Answer: D

Topic: Concept 2.4

Skill: Knowledge/Comprehension

72) Which of the following best describes chemical equilibrium?

- A) Forward and reverse reactions continue with no effect on the concentrations of the reactants and products.
- B) Concentrations of products are higher than the concentrations of the reactants.
- C) Forward and reverse reactions have stopped so that the concentration of the reactants equals the concentration of the products.
- D) Reactions stop only when all reactants have been converted to products.
- E) There are equal concentrations of reactants and products, and the reactions have stopped.

Answer: A

Topic: Concept 2.4

Skill: Knowledge/Comprehension

73) Which of the following describes any reaction that has reached chemical equilibrium?

- A) The concentration of the reactants equals the concentration of the products.
- B) The rate of the forward reaction is equal to the rate of the reverse reaction.
- C) All of the reactants have been converted to the products of the reaction.
- D) All of the products have been converted to the reactants of the reaction.
- E) Both the forward and the reverse reactions have stopped with no net effect on the concentration of the reactants and the products.

Answer: B

Topic: Concept 2.4

Skill: Knowledge/Comprehension

74) A group of molecular biologists is trying to synthesize a new artificial compound to mimic the effects of a known hormone that influences sexual behavior. They have turned to you for advice. Which of the following compounds is most likely to mimic the effects of the hormone?

- A) a compound with the same number of carbon atoms as the hormone

- B) a compound with the same molecular mass (measured in daltons) as the hormone
- C) a compound with the same three-dimensional shape as part of the hormone
- D) a compound with the same number of orbital electrons as the hormone
- E) a compound with the same number of hydrogen and nitrogen atoms as the hormone

Answer: C

Topic: Concept 2.4

Skill: Application/Analysis

Self-Quiz Questions

1) In the term trace element, the modifier trace means

- A) the element is required in very small amounts.
- B) the element can be used as a label to trace atoms through an organism's metabolism.
- C) the element is very rare on Earth.
- D) the element enhances health but is not essential for the organism's long-term survival.
- E) the element passes rapidly through the organism.

Answer: A

2) Compared with ^{31}P , the radioactive isotope ^{32}P has

- A) a different atomic number.
- B) one more neutron.
- C) one more proton.
- D) one more electron.
- E) a different charge.

Answer: B

3) Atoms can be represented by simply listing the number of protons, neutrons, and electrons for example, $2p^+$; $2n^0$; $2e^-$ for helium. Which one of the following lists represents the ^{18}O isotope of oxygen?

- A) $6p^+$; $8n^0$; $6e^-$
- B) $8p^+$; $10n^0$; $8e^-$
- C) $9p^+$; $9n^0$; $9e^-$
- D) $7p^+$; $2n^0$; $9e^-$
- E) $10p^+$; $8n^0$; $9e^-$

Answer: B

4) The atomic number of sulfur is 16. Sulfur combines with hydrogen by covalent bonding to form a compound, hydrogen sulfide. Based on the number of valence electrons in a sulfur atom, predict the molecular formula of the compound:

- A) HS
- B) HS_2
- C) H_2S

- D) H_3S_2
- E) H_4S

Answer: C

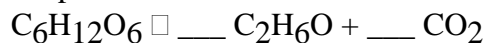
- 5) The reactivity of an atom arises from
- A) the average distance of the outermost electron shell from the nucleus.
 - B) the existence of unpaired electrons in the valence shell.
 - C) the sum of the potential energies of all the electron shells.
 - D) the potential energy of the valence shell.
 - E) the energy difference between the *s* and *p* orbitals.

Answer: B

- 6) Which statement is true of all atoms that are anions?
- A) The atom has more electrons than protons.
 - B) The atom has more protons than electrons.
 - C) The atom has fewer protons than does a neutral atom of the same element.
 - D) The atom has more neutrons than protons.
 - E) The net charge is 12.

Answer: A

- 7) What coefficients must be placed in the following blanks so that all atoms are accounted for in the products?



- A) 1; 2
- B) 2; 2
- C) 1; 3
- D) 1; 1
- E) 3; 1

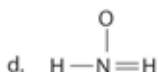
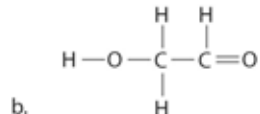
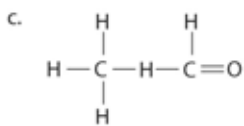
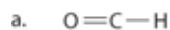
Answer: B

- 8) Which of the following statements correctly describes any chemical reaction that has reached equilibrium?

- A) The concentrations of products and reactants are equal.
- B) The rate of the forward reaction equals the rate of the reverse reaction.
- C) Both forward and reverse reactions have halted.
- D) The reaction is now irreversible.
- E) No reactants remain.

Answer: B

9) Draw Lewis structures for each hypothetical molecule shown below, using the correct number of valence electrons for each atom. Determine which molecule makes sense because each atom has a complete valence shell and each bond has the correct number of electrons. Explain what makes the other molecules nonsensical, considering the number of bonds each type of atom can make.



Answer:

a. $\ddot{\text{O}}::\overset{\uparrow}{\text{C}}::\text{H}$ This structure doesn't make sense because the valence shell of carbon is incomplete; carbon can form 4 bonds.

b.
$$\begin{array}{c} \text{H} \quad \text{H} \\ | \quad | \\ \text{H}:\ddot{\text{O}}:\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{C}}}\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{C}}}::\ddot{\text{O}} \\ | \\ \text{H} \end{array}$$
 This structure makes sense because all valence shells are complete, and all bonds have the correct number of electrons.

c.
$$\begin{array}{c} \text{H} \quad \text{H} \\ | \quad | \\ \text{H}:\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{C}}}\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{H}}}\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{C}}}\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{O}}} \\ | \\ \text{H} \end{array}$$
 This structure doesn't make sense because H has only 1 electron to share, so it cannot form bonds with 2 atoms.

d. This structure doesn't make sense for several reasons:

$$\begin{array}{c} \text{O} \\ | \\ \text{H}:\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{N}}}\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{H}}} \end{array}$$
 The valence shell of oxygen is incomplete; oxygen can form 2 bonds.
 H has only 1 electron to share, so it cannot form a double bond.

Nitrogen usually makes only 3 bonds. It does not have enough electrons to make 2 single bonds, make a double bond, and complete its valence shell.